Chapter 11 The Mole

Chapter 11 The Mole
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mole of water is the water molecule, the representative particle in a mole of copper is the copper atom, and the representative particle in a mole of sodium chloride is the formula unit. 310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02 1023 or one mole of representative particles. The representative particle for

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Chapter 11: The Mole 11.1 Measuring Matter counting particles, Avogadro's Number 6.02 x 10 23, mole, mole ↔ particle equations # moles x 6.022 x 10 23 particles/mole = # of particles # particles x 1 mole / 6.022 x 10 23 particles = # of moles. 11.2 Mass and the Mole

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mole cell chapter 11 1. Look at Figure 11.2 on page 310. Wha... 2. What is the term that is an SI unit... 3. In this book, Avogadro's number is r...

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the mole chapter 11 Flashcards and Study Sets | Quizlet 11.2 Mass and the Mole. Molar mass of an element. Equivalent to atomic mass. Unit is g/mol. Molar mass of C = Molar mass of Ra = Converting moles to mass. Ex] Calculate the mass in grams of 0.045 moles of chromium. Step 1 – Determine molar mass of Cr.

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11.2 Mass and the Mole. Molar mass of an element. Equivalent to atomic mass. Unit is g/mol. Molar mass of C = Molar mass of Ra = Converting moles to mass. Ex] Calculate the mass in grams of 0.045 moles of chromium. Step 1 – Determine molar mass of Cr. Step 2 – Use molar mass to convert known number of . moles to grams.

Chapter 11 – The Mole
In the rest of Chapter 11 we will use the mole to calculate the outcomes of reactions, how much of one substance reacts with another, how many molecules a mass of a substance contains, what are chemical reactions and how do we account for the number of molecules of one kind that react with another and how many different products molecules the reaction produces.

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The mole is the SI unit for the amount of a substance There are approximately 6.022 x 1023formula units in 1 mole of any substance or element This number is known as Avogadro's number

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Episode 11 – The Mole The World of Chemistry Episode 11 - The Mole 1. Why is it important to use the correct amount of materials in a chemical reaction7 2. What names are given to the materials at the beginning and end of a chemical reaction? 3. Atoms and molecules are extremely small. How do chemists "count" them? Can you think of

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The Mole. • The mass of a mole of a compound equals the sum of the masses of every particle that makes up the compound.The Molar Mass of Compounds A mole of a compound would contain

Avagadro's number of molecules of that compound. . The Molar Mass of Compounds • Ex. • • K 39.

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15.2 CHAPTER 11: STOICHIOMETRY. MOLE TO MOLE RATIO. When nitrogen and hydrogen gas are heated under the correct conditions, ammonia gas (NH3) is formed. a. RXN: 1. N2 + 3. H2 (2. NH3. b. How many moles of nitrogen react with three moles of hydrogen?___1 mol N2_____ 3 mol H2 1 mol N2. 3 mol H2. c.

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Chapter 10 Presentation. chapter 10: the mole. What is the mole? Chemists use the mole to count atoms, molecules, ions, and formula units. It is important to note that a mole always contains the same number of particles. For example, you know that a dozen eggs will always contain 12 eggs.

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